**Report from practicals**

**Volumetric analysis - titration**

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| Date of practical : | Group: | Choose the element. |
| Click here. | Student 1: | Click here to enter the text. |
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1. Explain why it is important to know the course of the titration curve when performing a titrimetric analysis.
2. Explain the basis for selecting the endpoint indicator for the titrimetric analysis method.
3. During acid-base analysis, the standard solution is added in small portions to the flask with the tested solution until ... (complete in your own words, describing this issue in detail).
4. During manganometric analysis, the standard solution is added in small portions to the flask with the tested solution until ... (complete in your own words, describing this issue in detail).
5. Calculate the percentage concentration of the tested solution. For the test, 2 ml of H2O2 solution was taken, added to a flask to which 5 ml of sulfuric acid solution and 50 ml of distilled water were also added, and then titrated with a standard KMnO4 solution with a concentration of 0.1 mol/l. Finally, 23.6 ml of the standard solution was used to obtain a pink color. 1 ml of this solution is equivalent to 1.7 mg of H2O2.